

Syn gas or Sabatier's Reaction

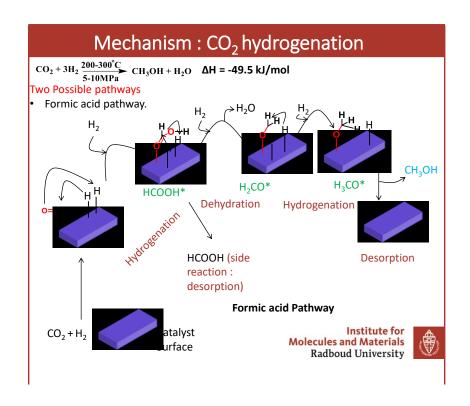
$$\begin{array}{c} \text{CO}_2 + 4 \text{H}_2 & \frac{450\text{-}600 \text{ K}}{5\text{-}10 \text{MPa}} & \text{CH}_4 + 2 \text{H}_2 \text{O} \\ \\ \text{CO}_2 + 3 \text{H}_2 & \frac{450\text{-}600 \text{ K}}{5\text{-}10 \text{MPa}} & \text{CH}_3 \text{OH} + \text{H}_2 \text{O} \\ \\ \text{CO}_2 + 2 \text{H}_2 & \frac{450\text{-}600 \text{ K}}{5\text{-}10 \text{MPa}} & \text{CH}_3 \text{OH} + \text{H}_2 \text{O} \\ \\ \text{CO}_3 + 2 \text{H}_2 & \frac{450\text{-}600 \text{ K}}{5\text{-}10 \text{MPa}} & \text{CH}_3 \text{OH} + \text{H}_2 \text{O} \\ \\ \text{CO}_4 + 2 \text{H}_2 & \frac{450\text{-}600 \text{ K}}{5\text{-}10 \text{MPa}} & \text{CH}_3 \text{OH} + \text{H}_2 \text{O} \\ \\ \text{CH}_3 + \text{H}_2 + \text{CH}_3 + \text{CH}_3$$

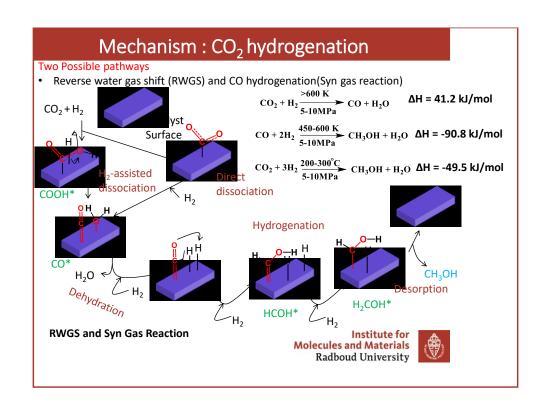
Metal Catalysts Used

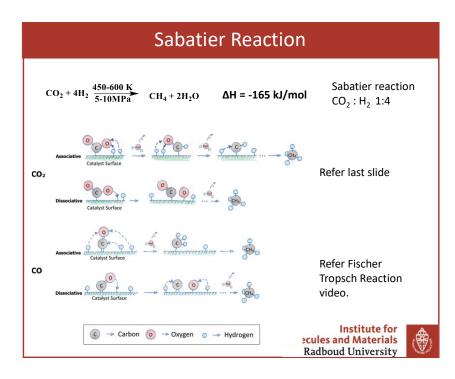
- Heterogeneous Catalysts.
- Non-scarce metals like nickel or copper on different supports like SiO₂, ZrO₂ etc.
- E.g. Cu/ZnO/Al₂O₃, Ni/ZrO₂











ICI process : Steam Reforming of Methane

- ICI: British Company
- Methane obtained from sources like coal mining, wastewater treatment, Sabatier process etc.
- Steam reforming or wet reforming is a reaction of methane and water at 1173K in presence of Ni/Al₂O₃ catalyst to form carbon monoxide and hydrogen in mole ratio 3:1.

$$CH_4 + H_2O \xrightarrow{1173 \text{ K}} CO + 3H_2$$

 Syn gas (CO + H₂) is hydrogenated to methanol by following pathway:

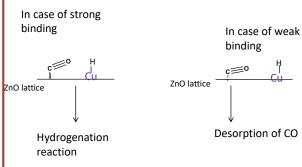
$$\begin{array}{ccc} CO + H_2 & & \underbrace{450\text{-}600 \text{ K}}_{5\text{-}10\text{MPa}} & CH_2O \\ \\ CH_2O + H_2 & \underbrace{450\text{-}600 \text{ K}}_{5\text{-}10\text{MPa}} & CH_3OH + H_2O \end{array}$$





Factors affecting reaction

- Deactivation of metal catalyst.
 - 4CO(g) + Ni(s) [Ni(CO)₄]
- Binding of CO with catalyst.







Methanol Utilization

· Methanol to dimethyl ether.

2CH₃OH
$$\longrightarrow$$
 CH₃OCH₃ + H₂O

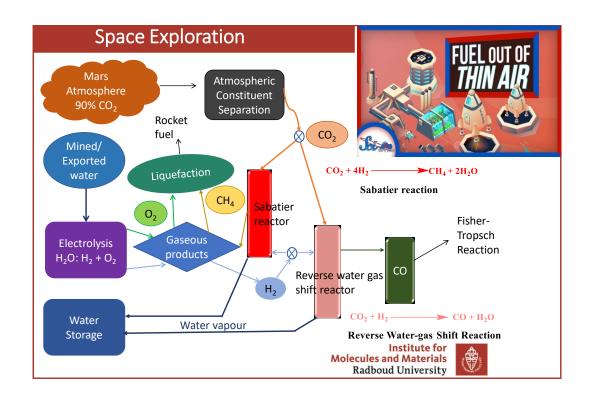
• Methanol utilization as a hydrogen carrier

Methanol steam reforming (MSR) is reaction of methanol and water at 500 K using Cu-based catalyst.

$$CH_3OH + H_2O \xrightarrow{500 \text{ K}} CO_2 + 3H_2$$

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Summary

- Methanol can be synthesized using Sabatier reaction which depends upon mole ratio of CO₂: H₂ or CO methanation using transition metal catalyst mostly nickel or copper.
- Mechanism can follow either formic acid or carbon monoxide pathway depending on catalyst.
- Industrially steam reforming of methane used.
- Methanol can be utilized as a hydrogen carrier: Methanol steam reforming (MSR) reaction.
- Sabatier reaction can let you live on mars.

